

- polar bond** is a bond between two different elements; due to the unequal sharing of the electrons, there is a dipole between the two atoms of the bond
- polar molecule** is one that has an overall dipole; it has dipoles associated with polar bonds that do not cancel as it lacks symmetry; examples are water and ammonia
- precipitation** is the formation of a solid during a chemical reaction when liquids or gases react
- reversible reactions** reactions where there is a forward reaction as well as a back reaction, e.g.. $N_2 + 3H_2 \rightleftharpoons 2NH_3$
- saturated solution** a solution in which no more solute can dissolve, excess solid settles on the bottom
- solute** the substance dissolved in a given solution; e.g.. sugar is the solute when it dissolves in water
- solution** when a gas, liquid or solid is dispersed homogeneously in a liquid; salt dissolved in water forms a solution
- solvent** a substance, present in the larger quantity, which dissolves a solute, forming a solution; e.g.. water is a solvent for sugar
- specific heat capacity** the amount of heat energy required to raise the temperature of 1 gram of the substance by 1°C (or by 1 K)
- surface tension** the force near the surface of a liquid due to unbalanced molecular forces; it causes the surface to assume a minimum area
- temperature scales** in Australia, as well as the majority of the world, temperature is measured in degrees Celsius; a minority of countries including the US, use Fahrenheit temperature; $^{\circ}C = 5/9(^{\circ}F - 32)$; absolute temperature, measured in kelvin, is used in scientific work, especially where calculation are involved
- thermal pollution** the release of heat into the environment, particularly pumping warm water from power plant cooling towers into rivers and lakes
- viscosity** a measure of a fluid's resistance to flow; e.g.. honey is more viscous than water, but petrol is less viscous
- weathering** the physical and chemical breakdown of rocks