

functional group

an atom or a group of atoms in all members of a homologous series that bestows certain chemical and physical properties onto the group

homologous

a series or family of similar carbon compounds differing in their number of  $\text{-CH}_2\text{-}$  groups but containing the same functional group

hydrolysis

a reaction with water

indicator

a substance that indicates when the concentration of a chemical species has passed a certain pH by a change in colour

ionisation reaction

the reaction between a molecular substance and water producing ions

IUPAC nomenclature	the system provided by the IUPAC for clearly naming chemicals with an explicit or implied relationship to the structure of compounds
Le Chatelier's principle	a principle that states that if a system at equilibrium is disturbed, the system tries to adjust itself so as to minimise that disturbance
neutralisation	the reaction between an acid and a base to produce salt and water only
neutral oxide	an oxide that displays neither acidic nor basic properties e.g. CO, N <sub>2</sub> O and NO
neutral salt	a substance formed when a strong acid is neutralised by a strong base or when a weak acid is neutralised by a weak base

primary standard

a substance of relatively high purity and stability that a solution of accurate concentration can be made from by direct weighing of a pure and dry chemical e.g. sodium carbonate

refluxing

a process of heating a reaction mixture in a vessel with an upright cooling condenser attached, preventing the loss of volatile reactants and products and allowing a higher temperature for the reaction

spectator ion

an ion present in solution and does not take part in the reaction, there to preserve charge neutrality

standard solution

a solution that has an accurately known concentration

strong acid

a solution in which the acid is effectively 100% ionised and/or dissociated

titrant	a solution of known composition and concentration used during titrations
titration	a common technique of volumetric analysis in which a standard solution of one reagent is added little by little from a burette to a second reagent whose concentration is to be determined until the end point is reached
volumetric analysis	a quantitative analysis of solutions having unknown concentration of some chemical, though the volume of the solution is known, by adding a reagent of known concentration until a reaction end point is reached
weak acid	a solution in which the acid is ionised only to a small extent