

1. <b>activity series of metals</b>	a tool which shows the relative reactivity of common metals from most reactive to least reactive, based on the chemical reactions they undergo	16. <b>law of conservation of matter</b>	a statement that matter can neither be created nor destroyed; it can only be changed from one form to another
2. <b>alloy</b>	a homogeneous mixture of a metal with one or more metals (or carbon) to give different properties e.g. steel and brass	17. <b>mineral</b>	a naturally occurring solid with a fixed chemical composition from which a metal or other material can be obtained
3. <b>anode</b>	the positive electrode in an electrolysis cell	18. <b>molar mass</b>	the mass in grams of one mole of a substance with units of grams per mole; calculated by adding the atomic weights of all atoms in the substance
4. <b>atom</b>	the smallest particle of matter that can take part in a chemical reaction; consists of a nucleus surrounded by electrons	19. <b>mole</b>	the amount of substance that contains the same number of particles as there are in exactly 12.00 grams of carbon-12
5. <b>atomic weight</b>	the average mass of the atoms present in a naturally occurring element relative to the mass of an atom of carbon-12 taken as exactly 12 as the standard	20. <b>ore</b>	a natural material obtained from the crust of the Earth that contains metals or other material
6. <b>Avogadro's law</b>	a statement that equal volumes of all gases at the same temperature and pressure contain equal numbers of particles	21. <b>percentage composition</b>	the percentage by mass of each element of a compound
7. <b>Avogadro's number</b>	the number of particles in one mole of any substance; equal to $6.022 \times 10^{23}$	22. <b>periodic table</b>	a table of the chemical elements in order of atomic number, arranged in rows and columns to illustrate periodic similarities and trends in physical and chemical properties
8. <b>cathode</b>	the negative electrode in an electrolysis cell	23. <b>theoretical yield</b>	the quantity of product predicted from the balanced chemical equation when known quantities of reactants undergo reaction
9. <b>electrolysis</b>	the passing of a direct electric current through a solution or molten material to decompose it	24. <b>valency</b>	the combining power of an element
10. <b>electronegativity</b>	a measure of the ability of an element to attract electrons		
11. <b>empirical formula</b>	the formula for a compound representing its atomic or ionic composition expressed in simple whole numbers e.g. the empirical formula for benzene, C <sub>6</sub> H <sub>6</sub> is CH		
12. <b>half-equations</b>	an equation written to describe an oxidation or reduction half-reaction, showing the loss or gain of electrons by an atom, forming an ion		
13. <b>ionisation energy</b>	the energy required to remove an electron from an atom in the gas state		
14. <b>isotopes</b>	atoms with the same number of protons, but different numbers of neutrons and so different mass		
15. <b>law of combining volumes</b>	a statement that the volumes of reacting gases involved (at the same temperature and pressure) may be expressed in simple whole number ratios		