Chemistry 5 The Acidic Environment - Part 2

1.	**functional group**	an atom or a group of atoms in all members of a homologous series that bestows certain chemical and physical properties onto the group
2.	**homologous**	a series or family of similar carbon compounds differing in their number of -CH2- groups but containing the same functional group
3.	**hydrolysis**	a reaction with water
4.	**indicator**	a substance that indicates when the concentration of a chemical species has passed a certain pH by a change in colour
5.	**ionisation reaction**	the reaction between a molecular substance and water producing ions
6.	**IUPAC nomenclature**	the system provided by the IUPAC for clearly naming chemicals with an explicit or implied relationship to the structure of compounds
7.	**Le Chatelier's principle**	a principle that states that if a system at equilibrium is disturbed, the system tries to adjust itself so as to minimise that disturbance
8.	**neutralisation**	the reaction between an acid and a base to produce salt and water only
9.	**neutral oxide**	an oxide that displays neither acidic nor basic properties e.g. CO, N2O and NO
10.	**neutral salt**	a substance formed when a strong acid is neutralised by a strong base or when a weak acid is neutralised by a weak base
11.	**primary standard**	a substance of relatively high purity and stability that a solution of accurate concentration can be made from by direct weighing of a pure and dry chemical e.g. sodium carbonate
12.	**refluxing**	a process of heating a reaction mixture in a vessel with an upright cooling condenser attached, preventing the loss of volatile reactants and products and allowing a higher temperature for the reaction
13.	**spectator ion**	an ion present in solution and does not take part in the reaction, there to preserve charge neutrality
14.	**standard solution**	a solution that has an accurately known concentration
15.	**strong acid**	a solution in which the acid is effectively 100% ionised and/or dissociated
16.	**titrant**	a solution of known composition and concentration used during titrations
17.	**titration**	a common technique of volumetric analysis in which a standard solution of one reagent is added little by little from a burette to a second reagent whose concentration is to be determined until the end point is reached
18.	**volumetric analysis**	a quantitative analysis of solutions having unknown concentration of some chemical, though the volume of the solution is known, by adding a reagent of known concentration until a reaction end point is reached
19.	**weak acid**	a solution in which the acid is ionised only to a small extent