

1. calorimetry	the measurement of heat changes associated with chemical reactions and physical processes	17. kelvin	the scientific temperature scale with absolute zero as 0 K; a change of one kelvin is the same as a change of one degree Celsius; $K = ^\circ C + 273$
2. cohesion	the attraction between the molecules of a liquid; it holds the particles of the liquid together	18. molarity	the number of moles of solute per litre of solution
3. density	the mass of a substance divided by its volume; units are g/cm^3 ; i.e.. lead has a density of $11.3 g/cm^3$	19. non-polar molecule	has no overall dipole; some molecules are non-polar if they are symmetrical so dipoles cancel, as in linear $O = C = O$; elements such as O_2 are non-polar as the atoms share the electrons equally and there is no dipole
4. dilution	in terms of a solution; refers to the addition of water to a solution to decrease concentration; the number of moles of a solute is unchanged	20. phase descriptors	or state symbols- (s), (l), (aq), (g), -in equations show the state of each substance
5. dipole	the unequal charge distribution between two atoms so that the atoms of a bond have a δ^+ and a δ^-		
6. dipole-dipole force	an attractive intermolecular force between the dipoles of neighbouring polar molecules		
7. dispersion force	a weak attractive force between molecules due to the attraction between negative electrons of one molecule and the positive nucleus of another molecule		
8. dissolution	the dissolving of a substance in a liquid		
9. dynamic equilibrium	an equilibrium where molecules are undergoing the forward as well as the back reaction, at the same rate; e.g.. liquid water and water vapour in a sealed container		
10. endothermic reaction	is where energy is absorbed from the surroundings		
11. enthalpy (ΔH)	is the heat content of a system; the total of all the kinetic and potential energies for one mole of a substance		
12. equilibrium	for a reversible reaction occurs when the rate of the forward reaction equals the rate of the back reaction		
13. exothermic reaction	is where energy is released to the surroundings		
14. hydrogen bonding	is the strongest of the intermolecular forces; the slightly positive charge on the hydrogen of one molecule is strongly attracted to the slightly negative charge on the F, O or N of another molecule		
15. intermolecular forces	the forces between molecules; e.g.. dispersion forces, dipole-dipole forces and hydrogen bonding		
16. joule	is the SI unit for energy, such as heat		