

1. acidic oxide	an oxide that shows acidic properties, but not basic properties	21. endothermic	a reaction where heat is taken in
2. acidic salt	a substance formed when a strong acid is neutralised by a weak base	22. end point	the point during a titration when the indicator changes colour, signalling that the reactants have completely reacted
3. acid rain	rain that has a higher concentration of hydrogen ions than pure water	23. equivalence point	the point reached during a titration when enough base has been added to neutralise the acid or when enough acid has been added to neutralise the base
4. alkali	a water soluble compound of the alkali metals (or ammonia) and acts as a strong base producing a high concentration of hydroxide ions in aqueous solution	24. esterification	a chemical reaction in which an organic acid chemically bonds with an alcohol, with the elimination of water
5. amphiprotic	a substance that can act as both a proton donor and a proton acceptor e.g. H ₂ O	25. exothermic	a reaction where heat is given out
6. amphoteric oxide	an oxide that shows both acidic properties or basic properties depending on the condition e.g. ZnO		
7. arrhenius acid	a substance that, in solution, can produce hydrogen ions		
8. arrhenius base	a substance that, in solution, can produce an hydroxide ion		
9. basic oxide	the oxide of a metal that displays basic properties, but not acidic properties		
10. basic salt	a substance formed when a weak acid is neutralised by a strong base		
11. buffer	a chemical substance that prevents any large changes to its pH if small amounts of acid or base are added		
12. carboxylic acids	a major class of organic compounds having the general formula RCOOH where R is an organic group		
13. concentrated	a solution containing a relatively large amount of solute		
14. condensed reaction	a reaction where one of the products is water e.g. esterification		
15. conjugate acid-base pairs	acid and base pairs where the base has one proton less than the acid		
16. dilute	a solution containing a relatively small amount of solute		
17. dissociation	the process where ions that already exist in an ionic compound are released when that substance dissolves		
18. dynamic equilibrium	the rate of forward reaction being the same as the rate of reverse reaction in a chemical reaction		
19. emollient	a softening agent		
20. emulsifying agent	a substance that can disperse one liquid in another immiscible one		