

polar bond

is a bond between two different elements; due to the unequal sharing of the electrons, there is a dipole between the two atoms of the bond

polar molecule

is one that has an overall dipole; it has dipoles associated with polar bonds that do not cancel as it lacks symmetry; examples are water and ammonia

precipitation

is the formation of a solid during a chemical reaction when liquids or gases react

reversible reactions

reactions where there is a forward reaction as well as a back reaction, e.g..  $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$

saturated solution

a solution in which no more solute can dissolve, excess solid settles on the bottom

solute	the substance dissolved in a given solution; e.g.. sugar is the solute when it dissolves in water
solution	when a gas, liquid or solid is dispersed homogeneously in a liquid; salt dissolved in water forms a solution
solvent	a substance, present in the larger quantity, which dissolves a solute, forming a solution; e.g.. water is a solvent for sugar
specific heat capacity	the amount of heat energy required to raise the temperature of 1 gram of the substance by $1\text{C}^{\circ}$ (or by 1 K)
surface tension	the force near the surface of a liquid due to unbalanced molecular forces; it causes the surface to assume a minimum area

temperature scales

in Australia, as well as the majority of the world, temperature is measured in degrees Celsius; a minority of countries including the US, use Fahrenheit temperature;  $^{\circ}\text{C} = 5/9(^{\circ}\text{F} - 32)$ ; absolute temperature, measured in kelvin, is used in scientific work, especially where calculation are involved

thermal pollution

the release of heat into the environment, particularly pumping warm water from power plant cooling towers into rivers and lakes

viscosity

a measure of a fluid's resistance to flow; e.g.. honey is more viscous than water, but petrol is less viscous

weathering

the physical and chemical breakdown of rocks