allotropes	different forms of an element; they may have different physical or chemical properties
atmosphere	the envelope of gas, vapour and aerosol particles surrounding the Earth, forming constituent in the environment of most forms of terrestrial life
atomic number	the number of protons in the nucleus of an atom, defining the chemical element
balanced equation	an equation using chemical symbols, having equal numbers of each atom on both sides
biosphere	the region of the Earth inhabited by living things, including air, land and water

chemical changes	changes that lead to a new substance being formed
compound	a substance formed when two or more chemical elements are chemically bonded together in the same ratio
covalent molecules	atoms linked by chemical bonds with sharing of electrons e.g. oxygen, carbon dioxide
covalent network substance	a substance with covalent bonds between atoms extending in a 3-dimensional network e.g. diamond, silicon oxide
decomposition	a chemical reaction when a compound splits up into elements or simpler compounds

electrolysis	the chemical reaction occurring when an electric current passes through a liquid; often used for obtaining pure elements
electron	an elementary particle of an atom, found in shells surrounding the nucleus
element	a substance composed of atoms of the same atomic number, incapable of being broken down to simpler substances displaying the same properties
empirical formula	a formula giving the proportions of the elements present in a compound but not the actual numbers or arrangement of atoms
gravimetric analysis	a set of methods for quantitatively determining a sample based on mass

hydrosphere	all the water of the Earth, in the oceans, rivers, lakes etc.
ion	an atom or group of atoms that has become electrically charged by the gain or loss of electrons e.g. Cl-, Na+
ionic compounds	a substance with attraction between positive and negative ions e.g. NaCl
ionic equations	chemical equations that show the formation of ions by the loss or gain of electrons
isotopes	atoms of the same element that have the same atomic number but different mass number i.e. they have the same number of protons, but a different number of neutrons